

CHEMICAL BONDS	
Curso: 3º ESO	Asignatura: Física y química
Objetivos contenido	Objetivos lingüísticos
- Uniones entre átomos: moléculas y cristales	-Relacionar definiciones con los términos en inglés. - Conocer vocabulario propio de las propiedades de las sustancias.
Criterios de evaluación de la asignatura	
2.9. Conocer cómo se unen los átomos para formar estructuras más complejas y explicar las propiedades de las agrupaciones resultantes.	
Criterios de evaluación de bilingüismo	
- Relacionar definiciones en inglés con sus correspondientes conceptos.	
Vocabulario	
Metallic - Metálico Ionic - Iónico Covalent - Covalente Conductivity - Conductividad Hardness - Dureza Malleability - Maleabilidad Solubility - Solubilidad	Conductor - Conductor Bond - Enlace Shared among - Compartido entre Outermost - más externo
Competencias básicas	
CL; CMCT; CAA	

Chemical bonds

Your friend has created a presentation on chemical bonds but they've mixed up their terms and their definitions.

a. Can you help match the terms to the correct definitions?

Term 1 Metallic		Definition A One or more electrons from one atom are attached to the other atom to complete their valence shell.	Definition B Atoms share electrons to complete their valence shell.
Term 2 Covalent	Term 3 Ionic		
Term 4 Hardness	Term 5 Conductivity	Definition C Stronger bonds make substances harder.	Definition D Weaker bonds in metals make them more malleable.
Term 6 Solubility	Term 7 Malleability	Definition E The presence of free electrons makes the substance a good conductor.	Definition F Weaker bonds and the presence of ions increases the solubility of a substance.
			Definition G Free electrons are shared among positively charged metal ions.

b. Can you add the correct word to each sentence?

outermost stability combine compounds element uncombined

1. A substance made of only one kind of atom is called an _____ and all the atoms of an element are identical.
2. The atoms of different elements _____ through chemical bonds to form more complex substances called _____.
3. Atoms are rarely found _____ in nature, apart from noble gases.
4. Atoms combine with each other to increase their _____ and gain, lose or share electrons to complete their _____ shell.